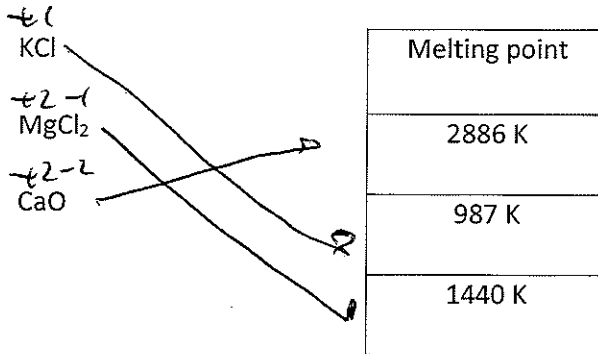


### #3-2 Properties of Ionic Substances/Melting Points

1. What are the criteria (in order) that you look for when determining the melting points of ionic substances?

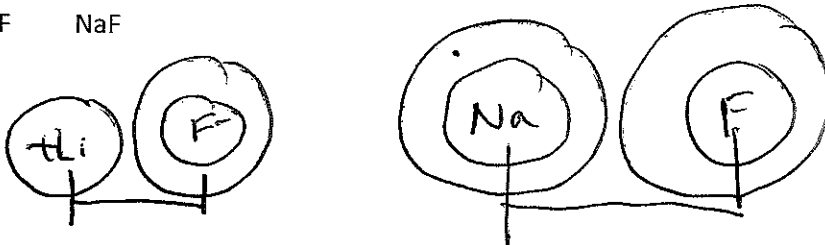
$F = \frac{q^+ q^-}{d^2}$   
 ← #1 charges  
 ← #2 inter nuclear distance

2. Match the chemical to its melting point.



3. Draw the Bohr diagrams of each substance.

LiF      NaF



Which has a higher melting point. Use Coulomb's Law to Justify.

LiF, due to a smaller inter nuclear distance

4. Arrange the following substances in order from lowest to highest melting points.

